Q1. A body is moving with constant speed, in a circle of radius 10 m. The body completes one revolution in 4 s. At the end of $3rd$ second, the displacement of body (in m) from its starting point is:

(1) 30 (2) $15\,\pi$

(3) $5\,\pi$ (4) $10\sqrt{2}$

Q2. Match List I with List II

	List I		**List II**
A	Angular momentum	I	$\left[\mathrm{ML}^2\,\mathrm{T}^{-2}\right]$
B	Torque	II	$\left[\mathrm{ML}^{-2}\,\mathrm{T}^{-2}\right]$
C	Stress	III	$\left[\mathrm{ML}^2\,\mathrm{T}^{-1}\right]$
D	Pressure gradient	IV	$\left[\mathrm{ML}^{-1}\,\mathrm{T}^{-2}\right]$

Choose the correct answer from the options given below :

(1) A-I, B-IV, C-III, D-II (2) A-III, B-I, C-IV, D-II

(3) A-II, B-III, C-IV, D-I (4) A-IV, B-II, C-I, D-III

Q3. A stone of mass 1 kg is tied to end of a massless string of length 1 m. If the breaking tension of the string is 400 N, then maximum linear velocity, the stone can have without breaking the string, while rotating in horizontal plane, is:

(1) $20\ \mathrm{m\ s^{-1}}$ (2) $40\ \mathrm{m\ s^{-1}}$

(3) $400\ \mathrm{m\ s^{-1}}$ (4) $10\ \mathrm{m\ s^{-1}}$

Q4. Two bodies are projected from ground with same speeds $40\ \mathrm{m\ s^{-1}}$ at two different angles with respect to horizontal. The bodies were found to have same range. If one of the body was projected at an angle of 60°, with horizontal then sum of the maximum heights, attained by the two projectiles, is _____ m. (Given $g = 10\ \mathrm{m\ s^{-2}}$)

Q5. A body of mass 10 kg is moving with an initial speed of $20\ \mathrm{m\ s^{-1}}$. The body stops after 5 s due to friction between body and the floor. The value of the coefficient of friction is: (Take acceleration due to gravity $g = 10\ \mathrm{m\ s^{-2}}$)

(1) 0.2 (2) 0.3

(3) 0.5 (4) 0.4

Q6. A ball is dropped from a height of 20 m. If the coefficient of restitution for the collision between ball and floor is 0.5, after hitting the floor, the ball rebounds to a height of _____ m.

Q7. Two discs of same mass and different radii are made of different materials such that their thicknesses are 1 cm and 0.5 cm respectively. The densities of materials are in the ratio 3 : 5. The moment of inertia of these discs respectively about their diameters will be in the ratio of $\frac{x}{6}$. The value of x is _____.

Q8. A body weight W, is projected vertically upwards from earth's surface to reach a height above the earth which is equal to nine times the radius of earth. The weight of the body at that height will be:

(1) $\frac{W}{91}$ (2) $\frac{W}{100}$

(3) $\frac{W}{9}$ (4) $\frac{W}{3}$

Q9. Under the same load, wire A having length 5.0 m and cross section $2.5 \times 10^{-5}\ \mathrm{m^2}$ stretches uniformly by the same amount as another wire B of length 6.0 m and a cross section of $3.0 \times 10^{-5}\ \mathrm{m^2}$ stretches. The ratio of the

Young's modulus of wire A to that of wire B will be:

(1) $1:4$ (2) $1:1$

(3) $1:10$ (4) $1:2$

Q10. A water heater of power 2000 W is used to heat water. The specific heat capacity of water is 4200 J kg^{-1} K^{-1}. The efficiency of heater is 70%. Time required to heat 2 kg of water from 10°C to 60°C is _____ s.
(Assume that the specific heat capacity of water remains constant over the temperature range of the water).

Q11. Heat energy of 735 J is given to a diatomic gas allowing the gas to expand at constant pressure. Each gas molecule rotates around an internal axis but do not oscillate. The increase in the internal energy of the gas will be:

(1) 525 J (2) 441 J

(3) 572 J (4) 735 J

Q12. A hypothetical gas expands adiabatically such that its volume changes from 08 litres to 27 litres. If the ratio of final pressure of the gas to initial pressure of the gas is $\frac{16}{81}$. Then the ratio of $\frac{C_p}{C_v}$ will be.

(1) $\frac{4}{3}$ (2) $\frac{3}{1}$

(3) $\frac{1}{2}$ (4) $\frac{3}{2}$

Q13. For a solid rod, the Young's modulus of elasticity is 3.2×10^{11} N m^{-2} and density is 8×10^3 kg m^{-3}. The velocity of longitudinal wave in the rod will be

(1) 145.75×10^3 m s^{-1} (2) 3.65×10^3 m s^{-1}

(3) 18.96×10^3 m s^{-1} (4) 6.32×10^3 m s^{-1}

Q14. The displacement equations of two interfering waves are given by $y_1 = 10 \sin\left(\omega t + \frac{\pi}{3}\right)$ cm, $y_2 = 5\left[\sin(\omega t) + \sqrt{3}\cos \omega t\right]$ cm respectively. The amplitude of the resultant wave is _____ cm.

Q15. Considering a group of positive charges, which of the following statements is correct?

(1) Net potential of the system cannot be zero at a point but net electric field can be zero at that point

(2) Net potential of the system at a point can be zero but net electric field can't be zero at that point

(3) Both the net potential and the net field can be zero at a point

(4) Both the net potential and the net electric field cannot be zero at a point

Q16. Two parallel plate capacitors C_1 and C_2 each having capacitance of 10 μF are individually charged by a 100 V D.C. source. Capacitor C_1 is kept connected to the source and a dielectric slab is inserted between it plates. Capacitor C_2 is disconnected from the source and then a dielectric slab is inserted in it. Afterwards the capacitor C_1 is also disconnected from the source and the two capacitors are finally connected in parallel combination. The common potential of the combination will be _____ V.
(Assuming Dielectric constant = 10)

Q17. The number of turns of the coil of a moving coil galvanometer is increased in order to increase current sensitivity by 50%. The percentage change in voltage sensitivity of the galvanometer will be :

(1) 100% (2) 50%

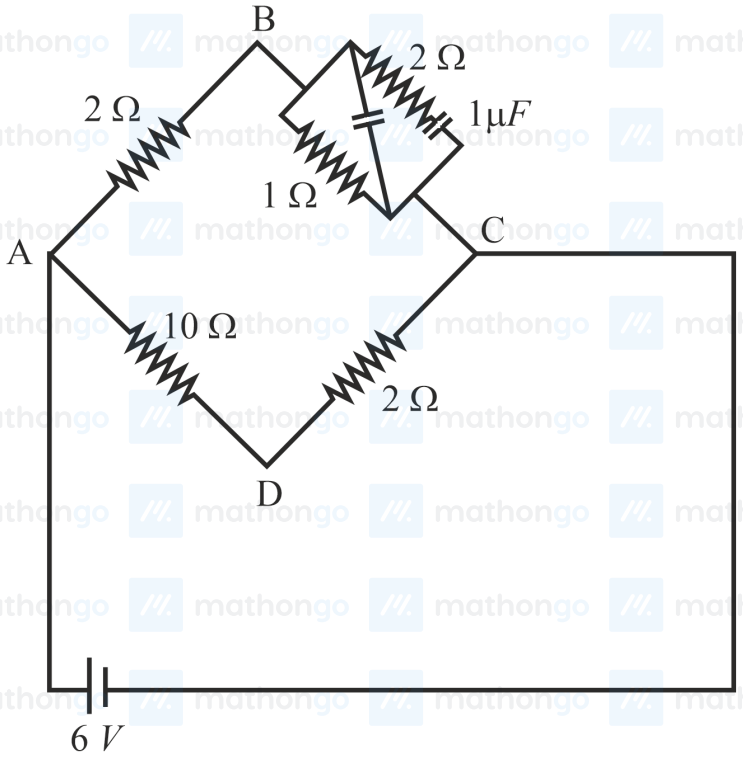
(3) 75% (4) 0%

Q18. The H amount of thermal energy is developed by a resistor in 10 s when a current of 4 A is passed through it. If the current is increased to 16 A, the thermal energy developed by the resistor in 10 s will be

(1) H (2) $16H$

(3) $\frac{H}{4}$ (4) $4H$

Q19. For the given circuit, in the steady state, $|V_B - V_D| =$ _____ V.



Q20. A long conducting wire having a current I flowing through it, is bent into a circular coil of N turns. Then it is bent into a circular coil of n turns. The magnetic field is calculated at the centre of coils in both the cases. The ratio of the magnetic field in first case to that of second case is:

(1) $N : n$ (2) $n^2 : N^2$

(3) $N^2 : n^2$ (4) $n : N$

Q21. An alternating voltage source $V = 260\sin(628t)$ is connected across a pure inductor of 5 mH. Inductive reactance in the circuit is:

(1) 3.14Ω (2) 6.28Ω

(3) 0.5Ω (4) 0.318Ω

Q22. A series LCR circuit consists of $R = 80\ \Omega$. $X_L = 100\ \Omega$, and $X_C = 40\ \Omega$. The input voltage is $2500\ \cos(100\pi t)$ V. The amplitude of current, in the circuit, is _____ A.

Q23. Match List I and List II

	List I		List II
A	Microwaves	I	Physiotherapy
B	UV rays	II	Treatment of cancer

C	Infra-red rays	III	Lasik eye surgery
D	X-rays	IV	Aircraft navigation

Choose the correct answer from the option given below:

(1) A-II, B-IV, C-III, D-I (2) A-IV, B-I, C-II, D-III

(3) A-IV, B-III, C-I, D-II (4) A-III, B-II, C-I, D-IV

Q24. A microscope is focused on an object at the bottom of a bucket. If liquid with refractive index $\frac{5}{3}$ is poured inside the bucket, then microscope have to be raised by 30 cm to focus the object again. The height of the liquid in the bucket is :

(1) 75 cm (2) 50 cm

(3) 18 cm (4) 12 cm

Q25. Two light waves of wavelengths 800 and 600 nm are used in Young's double slit experiment to obtain interference fringes on a screen placed 7 m away from plane of slits. If the two slits are separated by 0.35 mm , then shortest distance from the central bright maximum to the point where the bright fringes of the two wavelength coincide will be _____ mm.

Q26. If the two metals A and B are exposed to radiation of wavelength 350 nm. The work functions of metals A and B are 4.8 eV and 2.2 eV. Then choose the correct option

(1) Metal B will not emit photo-electrons (2) Both metals A and B will emit photo-electrons

(3) Both metals A and B will not emit photoelectrons (4) Metal A will not emit photo-electrons

Q27. The radius of electron's second stationary orbit in Bohr's atom is R. The radius of 3^{rd} orbit will be

(1) $\frac{R}{3}$ (2) $2.25\ R$

(3) $3R$ (4) $9R$

Q28. If the binding energy of ground state electron in a hydrogen atom is 13.6 eV, then, the energy required to remove the electron from the second excited state of Li^{2+} will be: $x \times 10^{-1}$ eV. The value of x is _____.

Q29. Given below are two statements:

Statement I : In a typical transistor, all three regions emitter, base and collector have same doping level.

Statement II : in a transistor, collector is the thickest and base is the thinnest segment.

In the light of the above statements, choose the most appropriate answer from the options given below.

(1) Both Statement I and Statement II are correct (2) Both Statement I and Statement II are incorrect

(3) Statement I is incorrect but Statement II is correct (4) Statement I is correct but Statement II is incorrect

Q30. Given below are two statements

Statement I : For transmitting a signal, size of antenna (l) should be comparable to wavelength of signal (at least $l = \frac{\lambda}{4}$ in dimension).

Statement II : In amplitude modulation, amplitude of carrier wave remains constant (unchanged).

In the light of the above statements, choose the most appropriate answer from the options given below

(1) Both Statement I and Statement II are correct
(2) Both Statement I and Statement II are incorrect
(3) Statement I is incorrect but Statement II is correct
(4) Statement I is correct but Statement II is incorrect

Q31. When a hydrocarbon A undergoes complete combustion it requires 11 equivalents of oxygen and produces 4 equivalents of water. What is the molecular formula of A?

(1) $C_{11}H_8$
(2) $C_{11}H_4$
(3) C_5H_8
(4) C_9H_8

Q32. Assume carbon burns according to following equation :

$2C(s) + O_2(g) \rightarrow 2\,CO\ (s)$

when 12 g carbon is burnt in 48 g of oxygen, the volume of carbon monoxide produced is $\_\_\_\_\_ \times 10^{-1}$ L at STP [nearest integer]

[Given : Assume CO as ideal gas, Mass of C is 12 g mol^{-1}, mass of *O* is 16 g mol^{-1} and molar volume of an idal gas at STP is 22.7 L mol^{-1}]

Q33. Arrange the following orbitals in decreasing order of energy.

A. $n = 3,\ l = 0,\ m = 0$
B. $n = 4,\ l = 0,\ m = 0$
C. $n = 3,\ l = 1,\ m = 0$
D. $n = 3,\ l = 2,\ m = 1$

The correct option for the order is:

(1) D > B > C > A
(2) B > D > C > A
(3) A > C > B > D
(4) D > B > A > C

Q34. Given below are two statements: one is labelled as Assertion (A) and the other is labelled as Reason (R).

Assertion (A): The first ionization enthalpy of 3d series elements is more than that of group 2 metals

Reason (R): In 3d series of elements successive filling of d-orbitals takes place.

In the light of the above statements, choose the correct answer from the options given below:

(1) Both (A) and (R) are true and (R) is the correct explanation of (A)
(2) (A) is true but (R) is false
(3) (A) is false but (R) is true
(4) Both (A) and (R) are true but (R) is not the correct explanation of (A)

Q35. Amongst the following, the number of species having the linear shape is

$XeF_2, I_3^+, C_3O_2, I_3^-, CO_2, SO_2, BeCl_2$ and $BCl_2^{\ominus}$

Q36. Enthalpies of formation of $CCl_4(g), H_2O(g), CO_2(g)$ and HCl are $-105, -242, -394$ and -92 kJ mol^{-1} respectively. The magnitude of enthalpy of the reaction given below is $kJmol^{-1}$. (nearest integer)

$CCl_4(g) + 2H_2O(g) \rightarrow CO_2(g) + 4\,HCl(g)$

Q37. Incorrect statement for the use of indicator in acid-base titration is:

(1) Methyl orange may be used for a weak acid vs weak bas titration.
(2) Phenolphthalein may be used for a strong acid vs strong base titration.
(3) Methyl orange is a suitable indicator for a strong acid vs weak base titration.
(4) Phenolphthalein is a suitable indicator for a weak acid vs strong base titration.

Q38. At 298 K, the solubility of silver chloride in water is 1.434×10^{-3} g L^{-1}. The value of $-\log K_{sp}$ for silver chloride is
(Given mass of Ag is 107.9 g mol^{-1}, and mass of Cl is 35.5 g mol^{-1})

Q39. Given below are two statements :
Statement I : H_2O_2 is used in the synthesis of Cephalosporin
Statement II : H_2O_2 is used for the restoration of aerobic conditions to sewage wastes.
In the light of the above statements, choose the most appropriate answer from the options given below :
(1) Statement I is correct but Statement II is incorrect
(2) Both Statement I and Statement II are correct
(3) Statement I is incorrect but Statement II is correct
(4) Both Statement I and Statement II are incorrect

Q40. The element playing significant role in neuromuscular function and interneuronal transmission is :
(1) Be
(2) Mg
(3) Ca
(4) Li

Q41. The number of alkali metal(s), from Li, K, Cs, Rb having ionization enthalpy greater than 400 kJ mol^{-1} and forming stable super oxide is

Q42. The Lewis acid character of boron tri halides follows the order:
(1) $BCl_3 > BF_3 > BBr_3 > Bl_3$
(2) $Bl_3 > BBr_3 > BCl_3 > BF_3$
(3) $BBr_3 > Bl_3 > BCl_3 > BF_3$
(4) $BF_3 > BCl_3 > BBr_3 > Bl_3$

Q43. In Dumas method for the estimation of N_2, the sample is heated with copper oxide and the gas evolved is passed over:
(1) Copper gauze
(2) Pd
(3) Ni
(4) Copper oxide

Q44. A hydrocarbon 'X' with formula C_6H_8 uses two moles of H_2 on catalystic hydrogenation of its one mole. On ozonolysis, 'X' yields two moles of methane dicarbaldehyde. The hydrocarbon 'X' is :
(1) hexa-1, 3, 5-triene
(2) 1-methylcyclopenta-1, 4−diene
(3) cyclohexa−1, 3−diene
(4) cyclohexa−1, 4−diene

Q45. The normal rain water is slightly acidic and its pH value is 5.6 because of which one of the following ?
(1) $CO_2 + H_2O \rightarrow H_2CO_3$
(2) $N_2O_5 + H_2O \rightarrow 2HNO_3$
(3) $2SO_2 + O_2 + 2H_2O \rightarrow 2H_2SO_4$
(4) $4NO_2 + O_2 + 2H_2O \rightarrow 4HNO_3$

Q46. A sample of a metal oxide has formula $M_{0.83}O_{1.00}$. The metal M can exist in two oxidation states +2 and +3. In the sample of $M_{0.83}O_{1.00}$, the percentage of metal ions existing in +2 oxidation state is %. (nearest integer)

Q47. Evaluate the following statements for their correctness.

A. The elevation in boiling point temperature of water will be same for 0.1 M NaCl and 0.1 M urea.

B. Azeotropic mixture boil without change in their composition.

C. Osmosis always takes place from hypertonic to hypotonic solution.

D. The density of 32% H_2SO_4 solution having molarity 4.09 M is approximately 1.26 g mL^{-1}.

E. A negatively charged sol is obtained when KI solution is added to silver nitrate solution.

Choose the correct answer from the options given below:

(1) B and D only (2) B, D and E only

(3) A and C only (4) A, B and C only

Q48. The resistivity of a 0.8M solution of an electrolyte is $5 \times 10^{-3}\Omega$cm. Its molar conductivity is $10^4\Omega^{-1}$ cm^2 mol^{-1}. (Nearest integer)

Q49. The rate constant for a first order reaction is 20 min^{-1}. The time required for the initial concentration of the reactant to reduce to its $\frac{1}{32}$ level is $\_\_\_\_\_ \times 10^{-2}$ min. (Nearest integer)

Q50. Match List I with List II

List I		**List II**	
A	Physisorption	I	Single Layer Adsorption
B	Chemisorption	II	$20-40$ kJ mol^{-1}
C	$N_2(g)+3H_2(g) \xrightarrow{Fe(s)} 2NH_3(g)$	III	Chromatography
D	Analytical Application or Adsorption	IV	Heterogeneous catalysis

Choose the correct answer from the options given below:

(1) A-III, B-IV, C-I, D-II (2) A-II, B-III, C-I, D-IV

(3) A-II, B-I, C-IV, D-III (4) A-IV, B-II, C-III, D-I

Q51. Which one of the following statements is incorrect ?

(1) van Arkel method is used to purify tungsten.

(2) The malleable iron is prepared from cast iron by oxidising impurities in a reverberatory furnace.

(3) Cast iron is obtained by melting pig iron with scrap iron and coke using hot air blast.

(4) Boron and Indium can be purified by zone refining method.

Q52. Which of the following elements have half-filled f-orbitals in their ground state ?

(Given : atomic number $Sm-62; Eu=63; Tb=65; Gd=64, Pm=61$)

Choose the correct answer from the options given below :

(1) A and B only (2) C and D only

(3) B and D only (4) A and E only

Q53. If the CFSE of $[Ti(H_2O)_6]^{3+}$ is -96.0 kJ / mol, this complex will absorb maximum at wavelength nm. (nearest integer)

Assume Planck' constant $(h) = 6.4 \times 10^{-34}$ Js, Speed of light $(c) = 3.0 \times 10^8$ m/s and Avogadro's constant $(N_A) = 6 \times 10^{23}$/ mol.

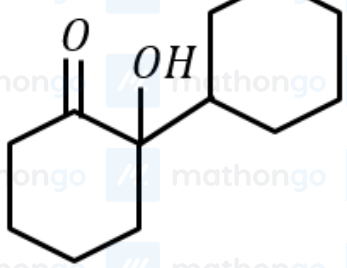
Q54. In the following halogenated organic compounds the one with maximum number of chlorine atoms in its structure is:

(1) Chloral (2) Gammaxene

(3) Chloropicrin (4) Freon-12

Q55. Cyclohexylamine when treated with nitrous acid yields (P). On treating (P) with PCC results in (Q). When (Q) is heated with dil. NaOH we get (R) The final product (R) is:

(1)



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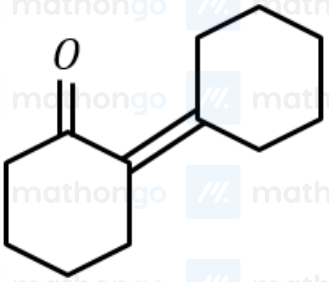

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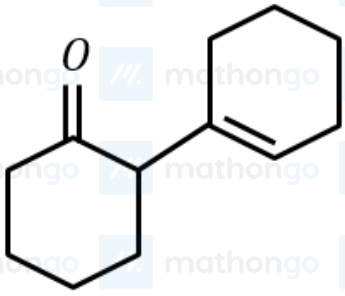

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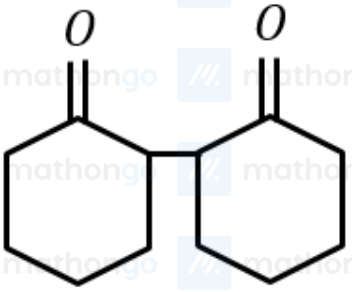
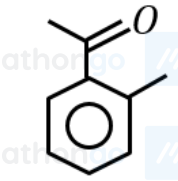
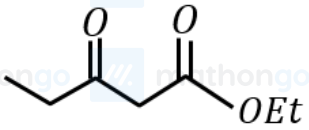
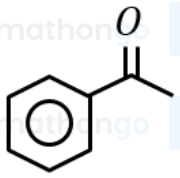
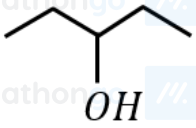


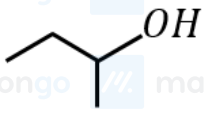
Q56. The number of molecules which gives haloform test among the following molecules is :

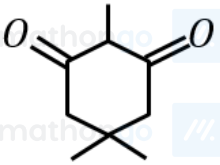




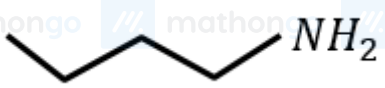


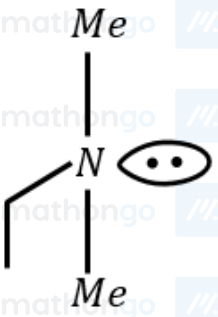


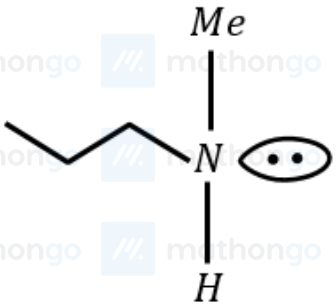


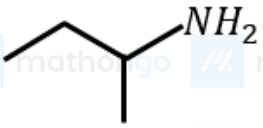


Q57. An organic compound [A]($C_4H_{11}N$), shows optical activity and gives N_2 gas on treatment with HNO_2. The compound [A] reacts with $PhSO_2Cl$ producing a compound which is soluble in KOH. The structure of A is :

(1) 

(2) 

(3) 

(4) 

Q58. Which of the following compounds are not used as disinfectants?

A. Chloroxylenol
B. Bithional
C. Veronal
D. Prontosil
E. Terpineol

Choose the correct answer from the options given below:

(1) A, B, E
(2) C, D
(3) A, B
(4) B, D, E

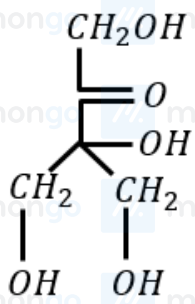
Q59. Given below are two statements :

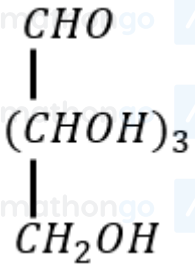
Statement I : Upon heating a borax bead dipped in cupric sulphate in a luminous flame, the colour of the bead becomes green.

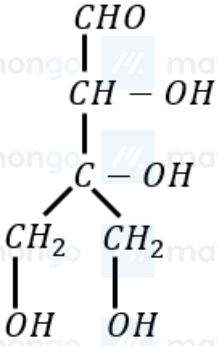
Statement II : The green colour observed is due to the formation of copper(I) metaborate In the light of the above statements, choose the most appropriate answer from the options given below:

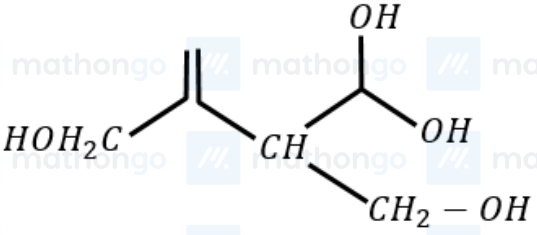
(1) Both Statement I and Statement II are false
(2) Both Statement I and Statement II are true
(3) Statement I is false but Statement II is true
(4) Statement I is true but Statement II is false

Q60. Compound A, $C_5H_{10}O_5$, given a tetraacetate with AC_2O and oxidation of A with $Br_2 - H_2O$ gives an acid, $C_5H_{10}O_6$. Reduction of A with HI gives isopentane. The possible structure of A is :

(1) 

(2) 

(3) 

(4) 

Q61. The equation $e^{4x}+8e^{3x}+13e^{2x}-8e^{x}+1=0, x \in R$ has :

(1) four solutions two of which are negative

(2) two solutions and both are negative

(3) no solution

(4) two solutions and only one of them is negative

Q62. The complex number $z=\frac{i-1}{\cos\frac{\pi}{3}+i\sin\frac{\pi}{3}}$ is equal to:

(1) $\sqrt{2}i\left(\cos\frac{5\pi}{12}-i\sin\frac{5\pi}{12}\right)$

(2) $\cos\frac{\pi}{12}-i\sin\frac{\pi}{12}$

(3) $\sqrt{2}\left(\cos\frac{\pi}{12}+i\sin\frac{\pi}{12}\right)$

(4) $\sqrt{2}\left(\cos\frac{5\pi}{12}+i\sin\frac{5\pi}{12}\right)$

Q63. Let $a_1, a_2, a_3, \ldots\ldots$ be an A.P. If $a_7=3$, the product (a_1a_4) is minimum and the sum of its first n terms is zero then $n!-4a_{n(n+2)}$ is equal to

(1) $\frac{381}{4}$

(2) 9

(3) $\frac{33}{4}$

(4) 24

Q64. The sum $1^2-2.3^2+3.5^2-4.7^2+5.9^2-\ldots\ldots+15.29^2$ is _____ .

Q65. The coefficient of x^{-6}, in the expansion of $\left(\frac{4x}{5}+\frac{5}{2x^2}\right)^9$, is

Q66. If the constant term in the binomial expansion of $\left(\frac{x^{\frac{5}{2}}}{2}-\frac{4}{x^l}\right)^9$ is -84 and the coefficient of x^{-3l} is $2^{\alpha}\beta$ where $\beta<0$ is an odd number, then $|\alpha l-\beta|$ is equal to _____ .

Q67. If ${}^{2n+1}P_{n-1} : {}^{2n-1}P_n = 11:21$, then n^2+n+15 is equal to :

Q68. The set of all values of a^2 for which the line $x+y=0$ bisects two distinct chords drawn from a point $P\left(\frac{1+a}{2}, \frac{1-a}{2}\right)$ on the circle $2x^2+2y^2-(1+a)x-(1-a)y=0$, is equal to :

(1) $(8,\infty)$

(2) $(0,4]$

(3) $(4,\infty)$

(4) $(2,12]$

Q69. Let S be the set of all $a \in N$ such that the area of the triangle formed by the tangent at the point $P(b, c), b, c \in N,$ on the parabola $y^2 = 2ax$ and the lines $x = b, y = 0$ is 16 unit^2, then $\sum_{a \in S} a$ is equal to _____ .

Q70. Let H be the hyperbola, whose foci are $\left(1 \pm \sqrt{2}, 0\right)$ and eccentricity is $\sqrt{2}$. Then the length of its latus rectum is:

(1) 3 (2) $\frac{5}{2}$
(3) 2 (4) $\frac{3}{2}$

Q71. $\lim\limits_{x \to \infty} \frac{\left(\sqrt{3x+1}+\sqrt{3x-1}\right)^6+\left(\sqrt{3x+1}-\sqrt{3x-1}\right)^6}{\left(x+\sqrt{x^2-1}\right)^6+\left(x-\sqrt{x^2-1}\right)^6} x^3$

(1) is equal to $\frac{27}{2}$ (2) is equal to 9
(3) does not exist (4) is equal to 27

Q72. The number of values of $r \in \{p, q, \sim p, \sim q\}$ for which $((p \wedge q) \Rightarrow (r \vee q)) \wedge ((p \wedge r) \Rightarrow q)$ is a tautology, is :

(1) 1 (2) 2
(3) 4 (4) 3

Q73. Let the mean and standard deviation of marks of class A of 100 students be respectively 40 and $\alpha(>0)$, and the mean and standard deviation of marks of class B of n students be respectively 55 and $30-\alpha$. If the mean and variance of the marks of the combined class of $100+n$ students are respectively 50 and 350 , then the sum of variances of classes A and B is

(1) 500 (2) 450
(3) 650 (4) 900

Q74. Among the relations
$S = \{(a, b) : a, b \in R - \{0\}, 2 + \frac{a}{b} > 0\}$ and $T = \{(a, b) : a, b \in R, a^2 - b^2 \in Z\}$,

(1) S is transitive but T is not (2) both S and T are symmetric
(3) neither S not T is transitive (4) T is symmetric but S is not

Q75. Let $A = \left[a_{ij}\right] \cdot a_{ij} \in Z \cap [0, 4],\ 1 \leq i, j \leq 2$. The number of matrices A such that the sum of all entries is a prime number $p \in (2, 13)$ is _____ .

Q76. Let A be a $n \times n$ matrix such that $|A| = 2$. If the determinant of the matrix $\text{Adj}\left(2.\text{Adj}\left(2\ \text{A}^{-1}\right)\right)$ is 2^{84}, then n is equal to _____ .

Q77.
If a point $P(\alpha, \beta, \gamma)$ satisfying $\begin{pmatrix}\alpha & \beta & \gamma\end{pmatrix}\begin{pmatrix}2 & 10 & 8 \\ 9 & 3 & 8 \\ 8 & 4 & 8\end{pmatrix} = \begin{pmatrix}0 & 0 & 0\end{pmatrix}$ lies on the plane $2x + 4y + 3z = 5$, then $6\alpha + 9\beta + 7\gamma$ is equal to

(1) $\frac{5}{4}$ (2) -1
(3) 11 (4) $\frac{11}{5}$

Q78. Let $(a, b) \subset (0, 2\pi)$ be the largest interval for which $\sin^{-1}(\sin\theta) - \cos^{-1}(\sin\theta) > 0, \theta \in (0, 2\pi)$, holds . If $\alpha x^2 + \beta x + \sin^{-1}\left(x^2 - 6x + 10\right) + \cos^{-1}\left(x^2 - 6x + 10\right) = 0$ and $\alpha - \beta = b - a$, then α is equal to;

(1) $\frac{\pi}{8}$ (2) $\frac{\pi}{48}$
(3) $\frac{\pi}{16}$ (4) $\frac{\pi}{12}$

Q79. Let $f : R - \{2, 6\} \to R$ be real valued function defined as $f(x) = \frac{x+2x+1}{x^2-8x+12}$. Then range of f is
(1) $\left(-\infty, -\frac{21}{4}\right] \cup \left[\frac{21}{4}, \infty\right)$ (2) $\left(-\infty, -\frac{21}{4}\right] \cup [0, \infty)$
(3) $\left(-\infty, -\frac{21}{4}\right) \cup (0, \infty)$ (4) $\left(-\infty, -\frac{21}{4}\right] \cup [1, \infty)$

Q80. The absolute minimum value, of the function $f(x) = |x^2 - x + 1| + [x^2 - x + 1]$, where $[t]$ denotes the greatest integer function, in the interval $[-1, 2]$, is
(1) $\frac{3}{2}$ (2) $\frac{1}{4}$
(3) $\frac{5}{4}$ (4) $\frac{3}{4}$

Q81. Let $\alpha > 0$. If $\int_0^{\alpha} \frac{x}{\sqrt{x+\alpha}-\sqrt{x}} dx = \frac{16+20\sqrt{2}}{15}$ then α is equal to :
(1) 2 (2) $2\sqrt{2}$
(3) 4 (4) $\sqrt{2}$

Q82. If $\phi\left(x\right) = \frac{1}{\sqrt{x}} \int_{\frac{\pi}{4}}^{x} \left(4\sqrt{2}\sin t - 3\phi'\left(t\right)\right) dt, x > 0$ then $\phi'\left(\frac{\pi}{4}\right)$ is equal to
(1) $\frac{4}{6+\sqrt{\pi}}$ (2) $\frac{8}{6+\sqrt{\pi}}$
(3) $\frac{8}{\sqrt{\pi}}$ (4) $\frac{4}{6-\sqrt{\pi}}$

Q83. Let the area of the region $\left\{(x, y) : |2x - 1| \leq y \leq \left|x^2 - x\right|, 0 \leq x \leq 1\right\}$ be A. Then $(6A + 11)^2$ is equal to _____ .

Q84. Let $y = y(x)$ be the solution of the differential equation $\left(3y^2 - 5x^2\right)ydx + 2x\left(x^2 - y^2\right)dy = 0$ such that $y(1) = 1$. Then $\left|(y(2))^3 - 12y(2)\right|$ is equal to :
(1) 64 (2) $32\sqrt{2}$
(3) 32 (4) $16\sqrt{2}$

Q85. Let $\vec{a} = \hat{i} + 2\hat{j} + 3\hat{k}, \vec{b} = \hat{i} - \hat{j} + 2\hat{k}$ and $\vec{c} = 5\hat{i} - 3\hat{j} + 3\hat{k}$, be there(three) vector. If $\vec{r}$ is a vector such that, $\vec{r} \times \vec{b} = \vec{c} \times \vec{b}$ and $\vec{r} \cdot \vec{a} = 0$, then $25\left|\vec{r}\right|^2$ is equal to
(1) 560 (2) 339
(3) 449 (4) 336

Q86. Let $\vec{a}, \vec{b}, \vec{c}$ be three vectors such that $\left|\vec{a}\right| = \sqrt{31}, 4\left|\vec{b}\right| = \left|\vec{c}\right| = 2$ and $2\left(\vec{a} \times \vec{b}\right) = 3\left(\vec{c} \times \vec{a}\right)$. If the angle between $\vec{b}$ and $\vec{c}$ is $\frac{2\pi}{3}$, then $\left(\frac{\vec{a} \times \vec{c}}{\vec{a} \cdot \vec{b}}\right)^2$ is equal to _____ .

Q87. Let the plane $P : 8x + \alpha_1 y + \alpha_2 z + 12 = 0$ be parallel to the line $L : \frac{x+2}{2} = \frac{y-3}{3} = \frac{z+4}{5}$. If the intercept of P on the y-axis is 1 , then the distance between P and L is
(1) $\sqrt{\frac{2}{7}}$ (2) $\frac{6}{\sqrt{14}}$
(3) $\sqrt{\frac{7}{2}}$ (4) $\sqrt{14}$

Q88. Let P be the plane, passing through the point $(1, -1, -5)$ and perpendicular to the line joining the points $(4, 1, -3)$ and $(2, 4, 3)$. Then the distance of P from the point $(3, -2, 2)$ is

(1) 6 (2) 4

(3) 5 (4) 7

Q89. The foot of perpendicular from the origin O to a plane P which meets the co-ordinate axes at the point A, B, C is $(2, a, 4), a \in N$. If the volume of the tetrahedron $OABC$ is 144 unit^3, then which of the following points is NOT on P?

(1) $(0, 4, 4)$ (2) $(3, 0, 4)$

(3) $(0, 6, 3)$ (4) $(2, 2, 4)$

Q90. Let A be the event that the absolute difference between two randomly chosen real numbers in the sample space $[0, 60]$ is less than or equal to a. If $P(A) = \frac{11}{36}$, then a is equal to _____ .

ANSWER KEYS

1. (4)	**2.** (2)	**3.** (1)	**4.** (4)	**5.** (2)	**6.** (2)	**7.** (1)	**8.** (1)
9. (4)	**10.** (1)	**11.** (4)	**12.** (2)	**13.** (3)	**14.** (1)	**15.** (3)	**16.** (1)
17. (4)	**18.** (2)	**19.** (3)	**20.** (4)	**21.** (80)	**22.** (5)	**23.** (5)	**24.** (300)
25. (20)	**26.** (55)	**27.** (1)	**28.** (25)	**29.** (48)	**30.** (136)	**31.** (4)	**32.** (1)
33. (1)	**34.** (1)	**35.** (2)	**36.** (3)	**37.** (2)	**38.** (1)	**39.** (4)	**40.** (1)
41. (1)	**42.** (3)	**43.** (1)	**44.** (3)	**45.** (2)	**46.** (2)	**47.** (4)	**48.** (2)
49. (1)	**50.** (3)	**51.** (227)	**52.** (5)	**53.** (173)	**54.** (10)	**55.** (2)	**56.** (59)
57. (25)	**58.** (17)	**59.** (480)	**60.** (3)	**61.** (2)	**62.** (4)	**63.** (4)	**64.** (1)
65. (3)	**66.** (4)	**67.** (2)	**68.** (1)	**69.** (4)	**70.** (3)	**71.** (4)	**72.** (2)
73. (4)	**74.** (1)	**75.** (2)	**76.** (2)	**77.** (2)	**78.** (4)	**79.** (3)	**80.** (2)
81. (6952)	**82.** (5040)	**83.** (98)	**84.** (45)	**85.** (146)	**86.** (204)	**87.** (5)	**88.** (125)
89. (3)	**90.** (10)						